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Josiah Willard Gibbs proposed the equation that scientists now use to calculate the free energy change for a process.  $\Delta G = \Delta H -$

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$T\Delta S$  When  $\Delta G$  is negative, the process is spontaneous and can do work (exergonic). When  $\Delta G$  is positive, the process is not spontaneous. Work must be done on the system to make it happen (endergonic).

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Students will be given some questions to answer before the lab is done. They will need to either rewrite the question or incorporate the question in the answer. The idea here is that

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when someone (like a college professor) looks at a student's lab notebook, they should be able to tell what the question was by merely looking at their lab report.

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