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produce a hydroxide ion in aqueous solution.

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-conductors of electricity... -can turn  
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-conductors of electricity... -feels slippery... -can....

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electrolytes in aqueous solution, and react with metals. Bases. Chapter 19 - Acids, Bases, and Salts Acid is hydrogen-ion donor ( $H^+$  or proton); base is hydrogen-ion acceptor. Acids

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vocabulary and major concepts presented in this chapter. compounds derived from the reaction of a strong base with a weak acid or a strong acid with a weak base or a solution of a weak base and one of its salts.

**Quia - Chapter 19 "Acids, Bases, and Salts"**

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634 Chapter 18 • Acids and Bases

Section 118.18.1 Figure 18.1

Rhododendrons flourish in rich, moist soil that is moderately acidic, whereas semper-vivum, commonly called hen and chicks, grows best in drier, slightly basic soil. Introduction to Acids and Bases-!).)DEADifferent models help describe the behavior of acids and

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bases.

## **Chapter 18: Acids and Bases**

strong acid? What is the pH of a  $1.5 \times 10^{-3}$  M solution of NaOH, a strong base? What is the molarity of a KOH solution if 25.0 mL of it is neutralized by 31.7 mL of a 0.100 M nitric acid solution? Write the steps in the complete ionization of the

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following polyprotic acids. 4.  $\text{H}_2\text{CO}_3$  5.  $\text{H}_3\text{BO}_3$  A solution has a  $[\text{H}^+]$  of  $1.0 \times 10^{-5}\text{M}$ . 17. lg. 19 ...

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Section 19.1 Acids-Bases Theories Pg.  
646-652 Hydronium ion ( $\text{H}_3\text{O}^+$  or  $\text{H}^+$ )  
- the ion  $\text{H}_3\text{O}^+$ , consisting of a



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protonated water molecule and present in all aqueous acids.

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